

Equilibrium constants for hydrolysis and associated equilibria in critical compilations

Germanium

Equilibrium reactions	lgK at infinite dilution and $T = 298 \text{ K}$		
	Baes and Mesmer, 1976	Wood and Samson, 2006	Filella and May, 2023
$\text{Ge(OH)}_4 \rightleftharpoons \text{GeO(OH)}_3^- + \text{H}^+$	-9.31	-9.32 ± 0.05	-9.099
$\text{Ge(OH)}_4 \rightleftharpoons \text{GeO}_2(\text{OH})_2^{2-} + 2 \text{H}^+$	-21.9		
$\text{GeO}_2(\text{OH})_2^{2-} + \text{H}^+ \rightleftharpoons \text{GeO(OH)}_3^-$			12.76
$8 \text{Ge(OH)}_4 \rightleftharpoons \text{Ge}_8\text{O}_{16}(\text{OH})_3^{3-} + 13 \text{H}_2\text{O} + 3 \text{H}^+$	-14.24		
$8 \text{Ge(OH)}_4 + 3 \text{OH}^- \rightleftharpoons \text{Ge}_8(\text{OH})_{35}^{3-}$			28.33
$\text{GeO}_2(\text{s, hexa}) + 2 \text{H}_2\text{O} \rightleftharpoons \text{Ge(OH)}_4$		-1.35	-1.373
$\text{GeO}_2(\text{s, tetra}) + 2 \text{H}_2\text{O} \rightleftharpoons \text{Ge(OH)}_4$	-4.37	-5.02	-4.999

C.F. Baes and R.E. Mesmer, *The Hydrolysis of Cations*. Wiley, New York, 1976, p. 349.

M. Filella and P.M. May, *The aqueous solution chemistry of germanium under conditions of environmental and biological interest: inorganic ligands*. *Applied Geochemistry*, 155, 105631 (2023).

doi:10.1016/j.apgeochem.2023.105631

S.A. Wood and I.M. Samson, *The aqueous geochemistry of gallium, germanium, indium and scandium*. *Ore Geol. Rev.*, 28, 57–102 (2006).

Distribution diagrams

These diagrams have been computed at two Ge concentrations ($1 \text{ mM} = 1 \times 10^{-3} \text{ mol L}^{-1}$ and $1 \text{ } \mu\text{M} = 1 \times 10^{-6} \text{ mol L}^{-1}$) with the 'best' equilibrium constants above (in green). Calculations assume $T = 298 \text{ K}$ for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).

