Radium

Equilibrium reaction	lg*K* at infinite dilution and *T* = 298 K	
	Nordstrom et al., 1990	Brown and Ekberg, 2016
$Ra^{2+} + H_2O \rightleftharpoons RaOH^+ + H^+$	−13.49	−13.49 ± 0.20

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 217−219.

D.K. Nordstrom, L.N. Plummer, D. Langmuir, E. Busenberg, H.M. May, B.F. Jones and D.L. Parkhurst, Revised chemical equilibrium data for major water-mineral reactions and their limitations. In: Chemical Modeling of Aqueous Systems II. D.C. Melchior and R.L. Bassett (eds.). ACS Symposium Series 416. ACS, Washington DC, 1990, pp. 398−446.

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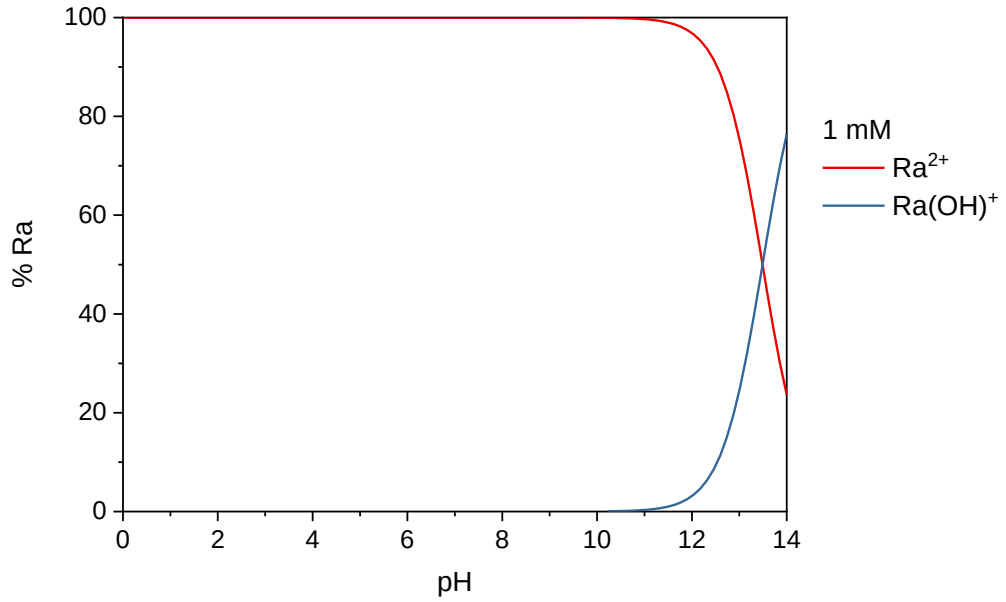
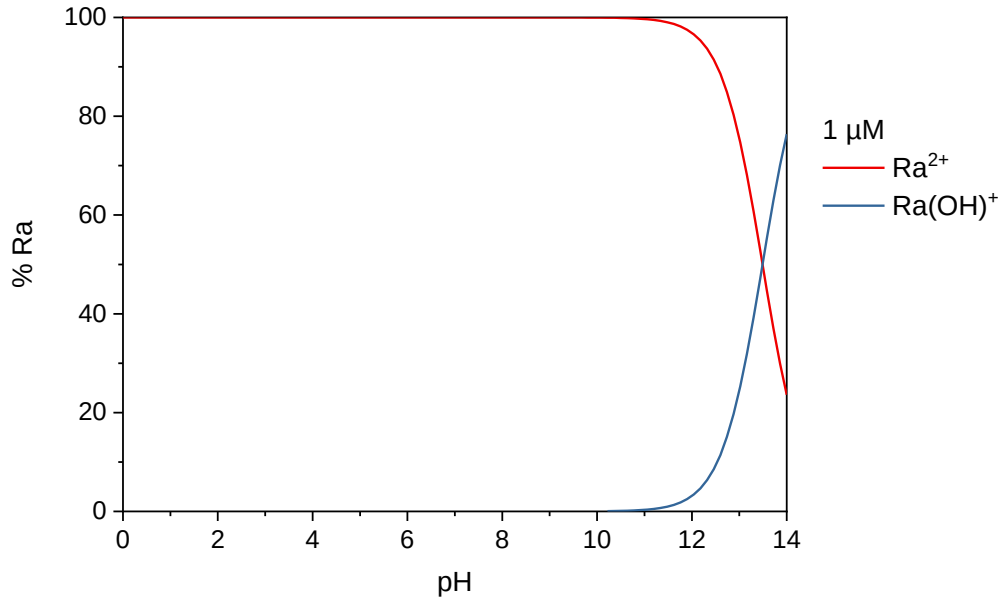
Distribution diagrams

These diagrams have been computed at two Ra concentrations (1 mM = $1x10^{-3}$ mol L^{-1} and 1 μM = $1x10^{-6}$ mol L^{-1}) with the 'best' equilibrium constant above (in green). Calculations assume T = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).





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